1366. The Crystal Structure of Tris-(o-diphenylarsinophenyl)arsineruthenium Dibromide

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A three-dimensional crystal structure analysis shows that the ruthenium(II) atom in a six-co-ordinated complex has bonds in an arrangement not greatly distorted from the octahedral in spite of stereochemical difficulties to be expected with the quadridentate ligand that it contains. This is made possible by considerable distortions of bond angles in the ligand.

THE ligand tris-(o-diphenylarsinophenyl)arsine [QAS = (I)] reacts with nitrosoruthenium compounds to form complexes of the type RuQASX₂ (X = Cl, Br, I, CNS, or NO₃) which are monomeric, diamagnetic, and non-electrolytes in nitrobenzene solution. So far as is known all finite hexa-co-ordinated complexes, including those of Ru^{II} are octahedral, but it seemed unlikely that an octahedral configuration could be preserved in the RuQASX₂ series, in view of the stereochemical limitations of the quadridentate ligand QAS. Indeed, were the ruthenium octahedrally co-ordinated then either the bonds between the ruthenium and the ligand or the bonds within the ligand itself must be considerably distorted and yet

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when X = Cl or Br the compound remains thermally stable at over 380° . To determine the environment of the ruthenium atoms in these compounds, the crystal structure of one member RuQASBr₂, has been examined.

Crystal Data—RuQASBr₂, $M = 1251 \cdot 5$; orthorhombic, $a = 31 \cdot 53 \pm 0.08$, $b = 11.40 \pm 0.10$, $c = 13.03 \pm 0.02$ Å; $U = 4684 \cdot 2$ Å³; $D_m = 1.79$ (by flotation); Z = 4, $D_c = 1.76$; Space group Pna2₁ (C⁹_{2v} No. 33); Cu-K_a radiation, single crystal photographs.

Structure Determination.—Over 2500 independent values of $F^2(hkl)$ were obtained from intensity measurements made from photographs taken by the equi-inclination Weissenberg method. Corrections were made for the Lorenz and polarisation factors but not for absorption or extinction. A three-dimensional Patterson map was computed from values of $F^2(hkl)$ sharpened to point atoms at rest.



the carbon atom referred to in the Tables.

The space groups $Pna2_1$ and Pnam are both compatible with the absent spectra, but since there are only four molecules in the unit cell Pnam would require the molecule to have a mirror plane of symmetry passing through the ruthenium atom. The position of the vectors could not be reconciled both with this requirement and with the known chemical formula. A trial structure based on $Pna2_1$ of four octahedra whose positions were determined from the vector map gave an R value of 0.38. Application of shifts, obtained from a difference map, to the positional and isotropic thermal parameters of the heavy atoms reduced R to 0.29. Two rounds of least-squares refinement on the positional and isotropic thermal parameters of the heavy atoms further reduced R to 0.20. The least-squares refinement employed the block diagonal approximation and a weighting scheme such that $1/w = 1 + [(F_o - b)/a]^2$ where a = 70 and b = 30.

At this stage a second difference synthesis showed the position of all 54 carbon atoms and two more rounds of least-squares refinement on the positional and isotropic thermal parameters of all except the hydrogen atoms reduced R to 0.087.

DISCUSSION

The purpose of the structure determination was to discover the environment of the heavy atoms and not to examine the detail of the phenyl groups. In the case of the carbon atoms the standard deviations are naturally higher than those in the positional parameters of the heavy atoms, but even so the bond angles and distances within the phenyl groups are not unreasonable.

Hartley and Venanzi ¹ reported that a molecular model showed that considerable ¹ Hartley and Venanzi, I_{\cdot} , 1962, 182.

distortion of the bond angles and distances would be necessary to form a complex in which the co-ordination of the ruthenium atom could be considered as approaching the octahedral. Nevertheless the co-ordinates of the heavy atoms establish that the bonds from ruthenium retain an octahedral configuration depiste the restraints and limitations imposed by the ligand.

The trigonal symmetry ² of the trigonal-bipyramidal ion [PtQASI)⁺ enables the bond angles at the apical arsenic atom As_1 to remain close to the tetrahedral, but it is not obvious how in the octahedral system the average deviation of the bonds at As_1 from the ideal tetrahedral can be as little as 2.5°, no greater than the average deviation (3.0°) in [PtQASI]⁺.

If the bonds from ruthenium were exactly octahedrally orientated and if the ruthenium atom were coplanar with each of the three bridging benzene rings, as is the platinum atom in [PtQASI]⁺, then the bond angles at As1 would not be as close to the tetrahedral as they are found to be and overcrowding would occur among the free phenyl groups.

The actual orientation found differs from the ideal octahedral in that the bonds from ruthenium to As3 and As4 are bent 7° towards Br1 and 7° towards As1 and the bond to As2 is bent 3° towards As1. The benzene ring forming the bridge between As1 and As2 is found to be coplanar with Ru, As1, As2, Br1, and Br2. The plane containing these atoms forms an approximate mirror plane of symmetry within the molecule. Because the bonds from As1 to As3 and As4 though they are coplanar with the arsenic atoms to which they are attached are not coplanar with ruthenium (see Figure). The potential overcrowding mentioned above is avoided partly by the effects of the distortions necessary to provide tetrahedral bonds at As1 and partly by the distortions of the bond angles at As2, As3, and As4 though the tetrahedral can be seen in Table 2. The largest deviations are found when bonds to non-bridging phenyls are involved. The occurrence of such deviations, which are geometrically impossible in the bridges, makes possible the retention of the octahedral form.

The bond from ruthenium to As1 is short compared with those to As2, As3, and As4. In the [PtQASI]⁺ ion the bond to As1 from platinum is also short ($2\cdot31$ Å) compared with those to As2 ($2\cdot49$ Å), As3 ($2\cdot45$ Å), and As4 ($2\cdot43$ Å). The shortening might be due to restrictions imposed by the ligand which in order to reduce the strain at As1 requires that the angles Ru-As1-C be slightly greater than the ideal tetrahedral.

The general conclusion from this structure is that the ruthenium bond angles are distorted from the octahedral about as much as they might be by the bridging requirements of a bidentate ligand. What seemed a possible way of preventing octahedral coordination fails because the valency angles at the arsenic atoms are not all sufficiently restricted by bridging.

EXPERIMENTAL

Preparation.—Samples were supplied by Dr. L. M. Venanzi, and had been recrystallised from tetrahydrofurfuryl alcohol. The yellow-orange crystals were about 0.25 mm. long.

X-Ray Photography.—The unit-cell dimensions were measured from zero layer Weissenberg photographs taken about the b and c axes. The intensities were estimated visually by the multiple film technique from equi-inclination Weissenberg photographs taken about the b axis. The layer scale factors were obtained by comparison of $F^{2}(hkO)$ values with those observed on a [001] zero layer Weissenberg photograph.

Calculations.—The main programmes used were written by J. S. Rollett (structure factor calculations and least-squares refinement), O. S. Mills (Fourier summations), G. A. Mair (standard deviations by procedures of Cruickshank³ and Darlow⁴). Final structure factors are

- ² Mair, Powell, and Venanzi, Proc. Chem. Soc., 1961, 170.
- ³ Ahmed and Cruickshank, Acta Cryst., 1953, 6, 385.
- ⁴ Darlow, Acta Cryst., 1960, 13, 683.

TABLE 1

Structure factors in absolute units with $(90 - \alpha)$ in degrees, given in every sixth column

k	h	l	F_{o}	$F_{\mathbf{c}}$		k	h	l	F_{o}	$F_{\mathbf{c}}$		k	h	l	F_{o}	F_{c}		k	h	l	F_{0}	$F_{\mathbf{c}}$	
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		4 56 7 8 9	42 70 60 29 35 27	36 70 57 30 42 28	161 14 213 32 129 38	3	2	2 3 4 56 9	22 63 36 43 28	22 54 28 43 34 25	208 14 170 219 98 101	3	II	0 I 2 3 4 5	106 302 147 240 267 183	101 310 139 243 268 189	90 171 79 178 145 175	3	2 I	2 3 4 5 0 1	19 20 20 36 34 19	17 33 26 39 28 24	1 38 1 4 1 28 3 0 3 2 7 0 3 2 9
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k	h	l	F_{o}	F_{c}		k	h	l	F_{0}	$F_{\mathbf{c}}$		k	h	l	F_{o}	F_{c}		k	h	l	$\boldsymbol{F}_{\mathrm{o}}$	F_{c}	
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3	33	8 0 3 4 5 6	22 71 38 32 56 33	27 78 43 36 58 47	39 270 263 277 256 271	4	8	11 12 13 0 1	22 21 28 131 87 91	29 20 41 131 87 85	94 261 42 270 19	4	17	11 12 0 1 2	21 19 24 80 35 18	26 28 23 80 40	22 237 90 55 250 203	4	27	4 5 0 1 2 3	31 38 114 22 58 31	34 42 119 23 62 30	239 183 90 287 95 220
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		3456790	197 12 136 101 84 81	123 94 78 81	303 191 68 226 208			56 78 9 10	98 56 28 21 31	97 47 37 24 37	25 173 19 244 113			1 2 3 5 7 8	26 27 29 48 38	24 17 27 34 47 46	44 17 156 121 281	4	.,	2 3 5 6 7	22 37 65 20 41	34 43 74 26 45	139 11 349 197 15
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4	2	1 2 3 4 5 6 8	18 54 131 65 69 77	54 57 59 65 76	30 10 14 343 177 8			4 5 7 9 10	64 75 56 19 21 31	62 66 45 18 24 36	220 216 206 53 250 143			3 4 5 8 9	75 53 36 21 38 31	78 57 39 23 47 40	159 194 344 30 77 106	4	32 33	4 50 8 6 1 7	44 27 47 15 24 47 18	49 32 59 15 19 55	2 57 2 52 2 70 3 2 6 3 3 5 1 8 8 1 7 7
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		9 10 11 12 13	123 54 86 76 89 20	52 88 74 93 28	336 107 317 95 311	4	13	0 1 2 3 4 5	50 40 52 50 100 83	43 39 50 52 97 81	90 168 85 95 149 109	4	2 I	10 0 1 2 3 4	20 86 63 67 29 30	33 85 67 62 40 29	73 270 223 107 149 133	4 4 5	36 37 0	1 1 4 3 5	17 15 14 16 123 188	23 14 16 33 116 186	327 305 2 304 183
4	4	14 15 0 1 2 3	20 21 59 46 10 72	30 30 61 41 6 66	120 90 270 339 276	4	13	7 9 10 1 2	28 31 22 71 175 34	45 30 32 68 177 33	95 107 155 270 318 281			96 78 90	30 75 31 31 42 20	33 75 33 38 49 27	236 325 107 15 215	5	I	7 9 11 2 3 4	75 63 47 57 30 41	53 50 45 23 30	244 158 215 132 70
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k	h	l	F_{o}	F_{c}		k	h	l	F_{0}	$F_{\mathbf{c}}$		k	h	l	F_{o}	F_{c}		k	h	l	F_{0}	$F_{\mathbf{c}}$	
5	2	XI X 2 34 56 8	36 18 97 41 105 24 163 68	31 23 94 40 96 19 145	332 120 89 355 113 335 85			2 3 4 5 6 7 8 9	76 21 54 88 76 61 64	81 19 50 77 72 58 61	206 20 263 3 197 322 206	5	21	7 8 • 10 11 3 4	36 47 27 24 34 73 58	29 47 30 22 42 74 55	306 101 130 332 248 331 333	5 5	32 33	56716012	45 125 186 250 250	50 19 28 19 27 71 30	143 12 149 64 326 90 230
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5	4	7 9 10 11 1 2 3	27 30 20 29 59 154 82 65	32 21 24 29 50 165 73 50	302 39 121 308 349 354 230 222	5	13	• 8 9 11 0 1 2 3	75 20 51 54 48 79 128 65	71 30 50 51 52 74 132 63	259 102 27 358 270 106 252 165			2 3 4 56 78 9	81 68 66 21 81 21 20 19	71 61 57 17 81 31 82	2 54 3 53 2 76 2 77 2 54 3 40 1 38 3 40	2	36 37	2 3 4 5 1 2 1 2	15 20 29 16 19 13 16 238	22 23 37 19 26 18 22 251	107 19 15 1 306 158 34 94
		4 56 7 8 9	87 155 32 58 33 35 36	82 146 30 50 31 37 34	346 3 111 340 311 39 45			4 56 7 8 9 10	93 18 62 56 29 21 30	76 25 62 57 39 20 27	293 160 233 83 222 179 294	5	23	0 2 56 78 9	50 21 47 21 23 32	41 29 44 23 25 34 38	90 212 221 88 209 233 211	6	z	46812456	183 233 64 42 13 18 42	181 222 53 51 7 21 39	97 83 228 353 110 217 292
5	2	11 0 1 2 3 4 5 6 7	55 76 58 79 75 88 75 88 75 88	52 81 75 48 71 31 66 80 14	350 270 126 60 248 169 230 256 125	3	14	2 3 4 5 6 7 8 9 0	57 66 99 55 51 53 21 36 75	54 60 86 45 46 . 41 38 29 74	175 16 330 57 215 74 274 43 201	.2	24	11 0 1 2 3 4 5 6 7	33 21 93 72 21 36 73 29 20	39 15 87 70 18 45 74 28 22	185 270 171 271 215 236 168 224 184	6	2	1 2 3 4 5 6 7 8	10 263 62 148 107 228 90 76 72	17 339 62 142 107 224 82 66 73	210 360 336 13 7 3 13 8 346
5	6	8 10 11 0 1 2 3 4	47 30 30 160 24 15 55 88	40 32 31 163 22 25 52 74	345 256 138 270 124 219 203 297	5	12	11 0 1 2 3 4 56	28 64 127 44 122 40 70 28	29 64 129 41 124 33 66 23	56 90 315 175 122 193 58	2	2 5	8 9 10 11 5 6	34 26 16 20 21 39 21 35	40 20 17 27 29 40 35 40	217 275 194 164 90 3 327 83	6	3	9 10 11 13 1 2 3 4	54 72 55 25 11 35 26 33	46 68 45 32 14 28 19 34	1 32 11 318 167 6 42 353 197
5	7	56 90 1 0 1 2 7	88 109 54 59 36 31 41 59	73 105 55 54 39 28 42 59	296 239 251 277 308 90 153 35 246	5	16	7 9 10 1 2 3 4 5 6	41 30 29 80 4 2 18 76 66	51 28 30 36 36 30 80 80	85 167 138 120 304 230 321 216 357	5	26	7 8 9 10 11 0 7 3	20 27 24 31 22 73 36 66 21	21 30 25 39 24 68 37 64	246 129 203 94 275 90 165 144	6	4	560123456	I 5 23 224 12 92 135 117 184 157	15 30 272 20 88 136 114 189	222 249 270 266 256 290 278 260
		34 56 78 90 I	69 72 29 63 47 20 47 66	5 5 2 2 5 8 2 9 2 9 2 9 4 7 5 9	134 184 45 180 21 230 126 178	5	17	78 90 11 0 1 2	62 59 47 54 33 83 17 119	53 60 48 49 43 89 19 122	134 289 207 341 16 2 90 247 67	5	27	+56 78 90 10	4 I 53 33 18 29 I 4 42 2 I	45 54 33 26 32 46 28	113 91 122 184 106 103 270 21	6	5	7 8 9 10 11 12 0 1	119 39 104 72 71 28 30 19	118 40 99 62 69 34 36 25	281 345 259 288 294 235 270 131
5	8	0 I 2 34 56 7	41 97 67 55 117 78 70 69	47 103 64 44 109 73 61 73	270 118 357 208 159 184 235 85	5	18	3 4 9 11 0 1 2	31 92 21 36 27 91 18 58	34 85 23 37 31 92 14 53	135 120 22 86 22 90 107 73			2 4 56 7 8 9 10	2 I 2 I 40 48 32 24 23 18	28 20 42 51 32 32 18 21	211 190 346 205 2 125 12 161	6	6	4 5 7 10 0 1 2 3	44 15 18 21 14 50 57 58	42 16 22 20 26 48 47 50	214 172 238 127 270 121 135 301
5	9	9 10 11 2 3 4	71 21 36 33 112 58 63	58 29 46 30 103 17 57 54	216 161 152 90 287 300 338 110			3 4 5 6 7 8 9 10	92 4 7 4 5 5 5 5 5 5 5 5 5 5 5 5 5	88 82 60 77 41 22 43 50	197 117 272 59 202 46 219 116	5	28	0 2 3 4 5 6 7 8	42 21 35 39 27 18 16	41 23 42 39 44 31 18 25	90 71 150 18 19 57 58 27			4 56 7 8 9 10 11	178 113 147 69 99 77 93 84	176 107 143 69 93 79 92 88	109 201 204 167 185 208 182 173
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		4	47	43	I 57			3	40	44	186	-		3	47	48	224						

TABLE 2

 $RuQASBr_2$

Some distances in Å and standard deviations in $Å \times 10^3$ shown in parentheses Within the bridging groups

	Within the br	raging groups		
2.308(5)	As3-C14	1.974(30)	C7-C8	1.438(57)
2.398(5)	As3-C31	1.945(40)	C8-C9	1.438(56)
$2 \cdot 465(5)$	As3-C37	1.901(40)	C9-C10	1.532(57)
$2 \cdot 472(5)$	As4-C8	1.947(42)	C10-C11	1·369(65)
2.610(5)	As4-C43	1.938(40)	C11-C12	1.369(55)
2.615(5)	As4-C49	1.991(40)	C12C7	1.413(47)
1.940(40)	C1-C2	$1 \cdot 400(47)$	C13-C14	1.398(51)
1.960(29)	C2-C3	1.441(52)	C14-C15	1.419(51)
1.947(39)	C3-C4	1.489(57)	C15-C16	1.417(52)
1.915(36)	C4-C5	$1 \cdot 365(51)$	C16-C17	1.458(58)
2.049(38)	C5-C6	$1 \cdot 420(59)$	C17-C18	1.369(57)
1.966(39)	C6-C1	1.448(60)	C18-C13	1.459(51)
	$\begin{array}{c} 2\cdot 308(5)\\ 2\cdot 398(5)\\ 2\cdot 465(5)\\ 2\cdot 472(5)\\ 2\cdot 610(5)\\ 2\cdot 615(5)\\ 1\cdot 940(40)\\ 1\cdot 960(29)\\ 1\cdot 947(39)\\ 1\cdot 915(36)\\ 2\cdot 049(38)\\ 1\cdot 966(39)\end{array}$	$\begin{array}{c cccc} & W \ thin \ the \ br \\ W \ thin \ the \ br \\ 2\cdot308(5) & As3-C14 \\ 2\cdot398(5) & As3-C31 \\ 2\cdot465(5) & As3-C37 \\ 2\cdot472(5) & As4-C8 \\ 2\cdot610(5) & As4-C43 \\ 2\cdot615(5) & As4-C43 \\ 2\cdot615(5) & As4-C49 \\ 1\cdot940(40) & C1-C2 \\ 1\cdot960(29) & C2-C3 \\ 1\cdot947(39) & C3-C4 \\ 1\cdot915(36) & C4-C5 \\ 2\cdot049(38) & C5-C6 \\ 1\cdot966(39) & C6-C1 \\ \end{array}$	$\begin{array}{c ccccccccccccccccccccccccccccccccccc$	$\begin{array}{c ccccccccccccccccccccccccccccccccccc$

In the phenyl groups

C-C Average 1.41; average deviation 0.04; maximum deviation 0.10 Å

Some angles in degrees and standard deviations in degrees \times 10 shown in parentheses

Some ang	sies in degrees a	nu stanuaru ueviatio	ns in degrees	× to shown in parent	lieses
As1-Ru-As2	86.7(2)	Ru-As2-C2	$106 \cdot 1(10)$	As2-C2-C1	121.5(28)
As1–Ru–As3	82.7(2)	Ru–As2–C19	$120 \cdot 3(27)$	As3-C14-C13	117.8(24)
Asl–Ru–As4	82.7(2)	Ru–As2–C25	120.7(11)	As4-C8-C7	121.9(28)
As2-Ru-As3	96·6(2)	C2-As2-C19	$108 \cdot 9(25)$	C6-C1-C2	$123 \cdot 2(36)$
As2–Ru–As4	97·1(2)	C2-As2-C25	150.5(15)	C1-C2-C3	116.4(33)
As3–Ru–As4	$159 \cdot 3(2)$	C19-As2-C25	94·2(36)	C2-C3-C4	120.9(30)
Brl-Ru-Asl	88·1(2)	Ru-As3-C14	103·5(9)	C3-C4-C5	118.0(37)
Br1-Ru-As2	$174 \cdot 8(2)$	Ru-As3-C31	$123 \cdot 8(11)$	C4C5C6	122.7(37)
Brl-Ru-As3	$82 \cdot 5(2)$	Ru–As3–C37	$123 \cdot 9(12)$	C5-C6-C1	117.3(35)
Brl-Ru-As4	$82 \cdot 5(2)$	C14-As3-C31	$97 \cdot 2(16)$	C12-C7-C8	121.7(31)
Brl-Ru-Br2	$92 \cdot 4(2)$	C14-As3-C37	104·9(15)	C7-C8-C9	120.1(35)
Br2–Ru–As1	178.9(2)	C31-As3-C37	99·2(16)	C8-C9-C10	115.3(36)
Br2–Ru–As2	92·8(2)	Ru-As4-C8	102 ·6(13)	C9-C10-C11	$119 \cdot 2(38)$
Br2–Ru–As3	98 ·4(2)	Ru-As4-C43	123·9(12)	C10-C11-C12	124.4(38)
Br2–Ru–As4	96·4(2)	Ru-As4-C49	$122 \cdot 3(11)$	C11-C12-C7	118.6(36)
Ru–Asl–Cl	110.3(11)	C8-As4-C43	$107 \cdot 2(17)$	C18-C13-C14	119.7(34)
Ru-Asl-C7	114.0(10)	C8-As4-C49	96·6(17)	C13-C14-C15	121.5(30)
Ru-Asl-C13	$111 \cdot 5(12)$	C43-As4-C49	100·2(16)	C14-C15-C16	118·4(34)
Cl-Asl-C7	104.0(15)	As1-C1-C2	$115 \cdot 3(28)$	C15-C16-C17	120.4(35)
Cl-Asl-Cl3	$109 \cdot 5(17)$	As1-C7-C8	$109 \cdot 2(24)$	C16-C17-C18	$120 \cdot 2(35)$
C7-As1-C13	$107 \cdot 1(13)$	As1–C13–C14	114.5(26)	C17-C18-C13	119.7(35)
	• •				

C-C-C in phenyl groups; average 119.9°; average deviation 3.9°; maximum deviation 10°.

TABLE 3

RuQASBr₂

Ato	mic co-ord	inates in fra	actions of th	ne unit c	ell lengths,	b, a, and	c and isot	ropic therm	al factors
in Å^2					-				
	Y	X	Ζ	В		Y	X	Ζ	B
Ru	0.17175	0.12016	0.00075	3.612	C25	0.3350	0.2004	-0.1634	4.245
Brl	0.02127	0.05987	0.03928	4.606	C26	0.2254	0.2143	-0.2056	9.166
Asl	0.31962	0.07205	0.03372	3.556	C27	0.2355	0.2429	-0.5884	9.655
As2	0.32378	0.17107	-0.03054	3.756	C28	0.3344	0.2541	-0.3448	5.363
Br2	0.00548	0.17469	-0.04045	4.587	C29	0.4313	0.2381	-0.3023	9.490
As3	0.17119	0.12685	0.18927	3.525	C30	0.4422	0.2092	-0.2172	8.916
As4	0.17533	0.08630	-0.17039	3.860	C31	0.2809	0.1620	0.2656	4.782
C1	0.4717	0.0983	0.0128	5.002	C32	0.3960	0.1473	0.2693	5.172
C2	0.4688	0.1411	-0.0121	3.365	C33	0.4818	0.1741	0.3259	5.047
C3	0.5798	0.1630	-0.0163	5.072	C34	0.4440	0.2131	0.3691	6.528
C4	0.6920	0.1392	-0.0063	7.237	C35	0.3283	0.2269	0.3618	6.050
C5	0.6863	0.0920	0.0164	5.478	C36	0.2462	0.2013	0.3034	6.555
C6	0.5789	0.0758	0.0369	7.388	C37	0.0339	0.1366	0.2697	3.892
C7	0.3208	0.0232	-0.0591	3.915	C38	0.0427	0.1428	0.3768	5.832
C8	0.2376	0.0302	-0.1404	5.316	C39	-0.0694	0.1532	0.4318	6.497
C9	0.2105	-0.0034	-0.2112	5.214	C40 -	-0.1694	0.1560	0.3806	7.544
C10	0.2874	-0.0431	-0.2021	7.063	C41 -	-0.1803	0.1474	0.2715	5.441
C11	0.3624	-0.0468	-0.1206	5.876	C42 -	-0.0736	0.1390	0.2121	4.798
C12	0.3842	-0.0152	-0.0510	5.257	C43	0.0378	0.0785	-0.2556	5.046
C13	0·3100	0.0200	0.1730	5.261	C44 -	-0.0660	0.0904	-0.2152	6.875
C14	0.2302	0.0713	0.2363	3.764	C45 ·	-0.1653	0.0823	-0.2721	5.823
C15	0.2049	0.0568	0.3369	5.093	C46 -	-0.1701	0.0683	-0.3702	7.540
C16	0.2635	0.0202	0.3734	5.670	C47 -	-0.0599	0.0550	-0.4153	5.600
C17	0.3464	-0.0023	0.3080	5.942	C48	0.0426	0.0623	-0.3556	6.083
C18	0.3680	0.0112	0.2104	5.016	C49	0.2931	0.1017	-0.2769	5.573
C19	0.3313	0.2259	0.0536	6.770	C50	0.4008	0.0864	-0.2635	6.868
C20	0.2269	0.2474	0.0520	6.112	C51	0.4850	0.1008	-0.3423	8.192
C21	0.2305	0.2867	0.1119	8 ∙380	C52	0.4606	0.1239	-0.4243	6.469
C22	0.3363	0.2986	0.1283	6.681	C53	0.3406	0.1389	-0.4344	6.757
C23	0.4306	0.2752	0.1622	5.911	C54	0.2588	0.1291	-0.3604	5.930
C24	0.4382	0.2356	0.1047	5.071					

given in Table 1, positional and thermal parameters in Table 3 and bond angles and distances together with the estimated standard deviations in Table 2.

We thank Dr. L. M. Venanzi for material, the D.S.I.R. for a maintenance grant (to R. H. B. M.), and the Director and staff of the Oxford University Computing Laboratory for facilities.

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